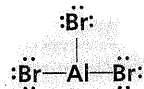
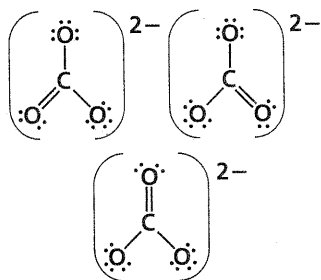
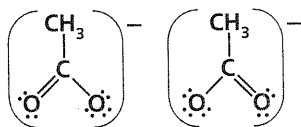
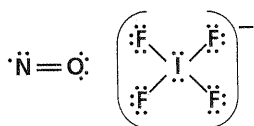


c. N_2O d. OCl_2 e. SiO_2 f. $AlBr_3$ 5. Draw the Lewis resonance structure for CO_3^{2-} .6. Draw the Lewis resonance structure for $CH_3CO_2^-$.7. Draw the Lewis structure for NO and IF_4^- .

Section 9-4

8. Determine the molecular geometry, bond angles, and hybrid of each molecule in question 4.

a. CO linear, 180° , sp b. CH_2O trigonal planar, 120° , sp^2 c. N_2O linear, 180° , sp d. OCl_2 bent, 104.5° , sp^3 e. SiO_2 linear, 180° , sp f. $AlBr_3$ trigonal planar, 120° , sp^2

Section 9-5

9. Determine whether each of the following molecules is polar or nonpolar.

a. CH_2O

The molecular shape is trigonal planar. The charge distribution is not symmetric due to the presence of more electronegative oxygen. The molecule is polar.

b. BF_3

The molecular shape is trigonal planar. The charge distribution is symmetric. The molecule is nonpolar.

c. SiH_4

The molecular shape is tetrahedral. The charge distribution is symmetric. The molecule is nonpolar.

d. H_2S

The molecular shape is bent. The charge distribution is not symmetric. The molecule is polar.