

Acids and Bases: Chapter 19

IV. Measuring strength

)...EQUATIONS FOR pH!

A. First, a little pHun (note: H<sup>+</sup>

1.

a. What is the pH of a 0.0100 M nitric acid solution, HNO<sub>3</sub>? Nitric acid is a strong acid.

b. pH=

2.

a. What is the pOH of a 0.050 M solution of NaOH?

b. pOH=

3. Therefore,

a. What is the hydronium (H<sub>3</sub>O<sup>+</sup>) concentration of a solution with a pH of 3.4?

b.

4. Relating pH and pOH

a.

b.

c.

d. What is the pH of a solution that has a pOH of 3.1?

e. What is the pOH of a solution that has a pH of 5.6?

