Teacher's Tools® Chemistry
Prerequisites: Concentration: Worksheet 4

1) What is the correct procedure for making 1.00 L of a 1.00 M solution of $CaCl_2$?

2) How many milliliters of 4.20-molar HBr must be diluted to obtain 1.0 liter of 1.0-molar HBr?

3) When 60.0 milliliter of 3.0-molar K_2CO_3 is added to 40.0 milliliters of 1.0-molar $KHCO_3$ the resulting concentration of K^+ is

4) The weight of H_2SO_4 (molar mass 98.1 g/mol) in 100.0 milliliters of a 6.00-molar solution is

5) What is the mass of NaCl required to prepare 40 mL of a 2.5 M solution?

6) How many moles of Cl^- are present in 70 mL of a 0.200 M $MgCl_2$ solution?

7) What is the molar concentration of ethanol in a solution made by mixing 23 g of ethanol with 200 mL of water? (density of ethanol = 0.789 g/mL and MW ethanol = 46 g/mol)

8) How many grams of calcium carbonate would be necessary to make 600 mL of a 2.0 M solution?