

# Kinetics - Chapter 17

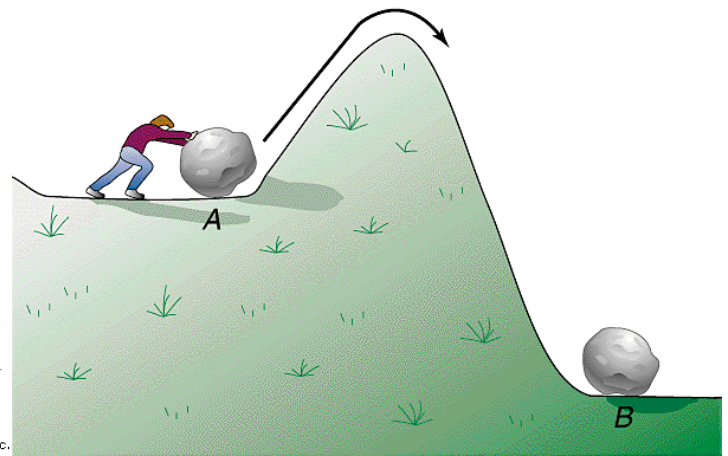
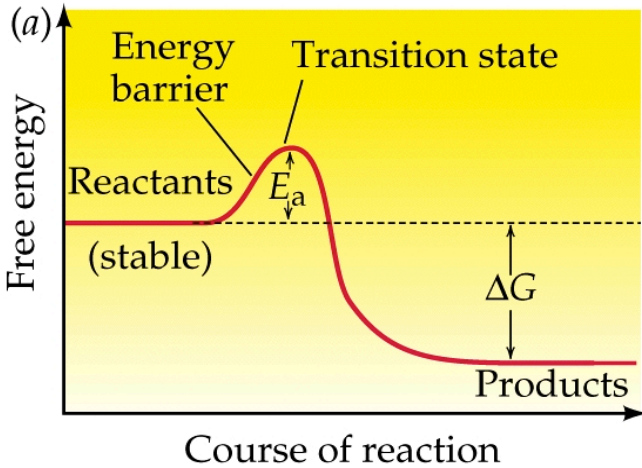
I.

- A. Chemical :
- B. :

1. Over time, the \_\_\_\_\_ and the \_\_\_\_\_
2. Reaction rate varies with the type of reaction
  - a. combustion = \_\_\_\_\_
  - b. rusting = \_\_\_\_\_

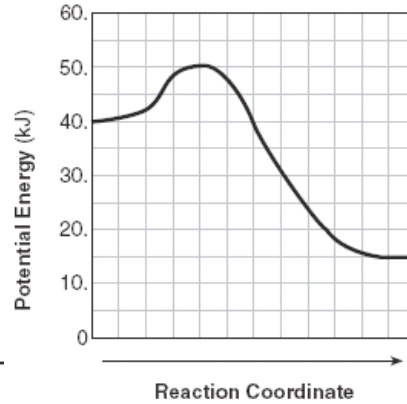
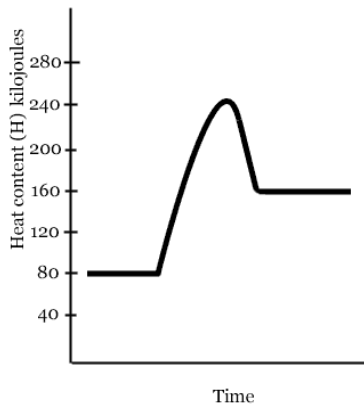
## II. The Reaction Process

- A. \_\_\_\_\_ :
1. Theory states that reactants must \_\_\_\_\_ for a reaction to occur.
  2. BUT,
  3. In order for an \_\_\_\_\_ (creation of new products) to occur molecules \_\_\_\_\_ with:
    - a. \_\_\_\_\_ (known as \_\_\_\_\_).
    - b. the \_\_\_\_\_
  4. The \_\_\_\_\_
  5. The minimum amount of \_\_\_\_\_ is called \_\_\_\_\_



Is this an endothermic or exothermic reaction?

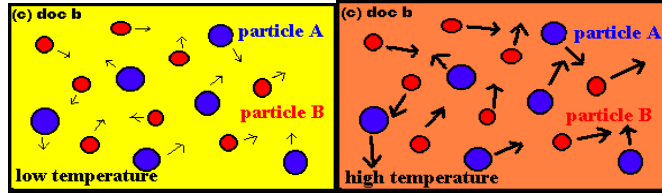
6. \_\_\_\_\_ : shows energy changes as a reaction proceeds.
  - a. When the \_\_\_\_\_
  - b. When the \_\_\_\_\_
  - c. The activation energy of any reaction is the \_\_\_\_\_



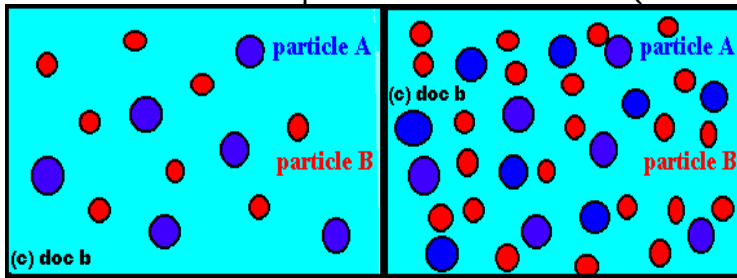
# Kinetics - Chapter 17

## III. Factors that Affect Reaction Rates

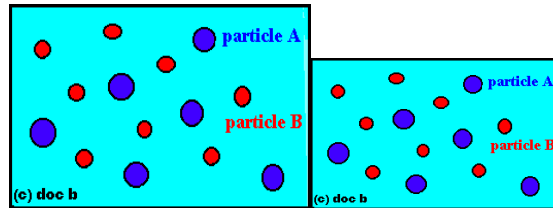
- A. : how complex and reactive the reactants are.  
 B. :  
 1. as the  
 2. This is due to an  
 3. Moving = (two birds with one stone!)



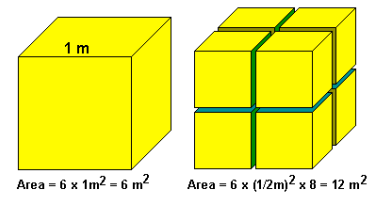
- C. :  
 1. as the  
 2. = for particles (but )!



- D. :  
 1. Rate increases as the pressure of reactants increases.  
 2. When pressure increases,  
 3. If you = (again )!  
 4. This is only true for



- E. :  
 1. An will  
 2. =



- F. :  
 1.  
 2. in the reaction  
 3. Catalysts by  
 4. In order to do this, the catalysts create a different pathway.  
 5. are the body's catalysts!

