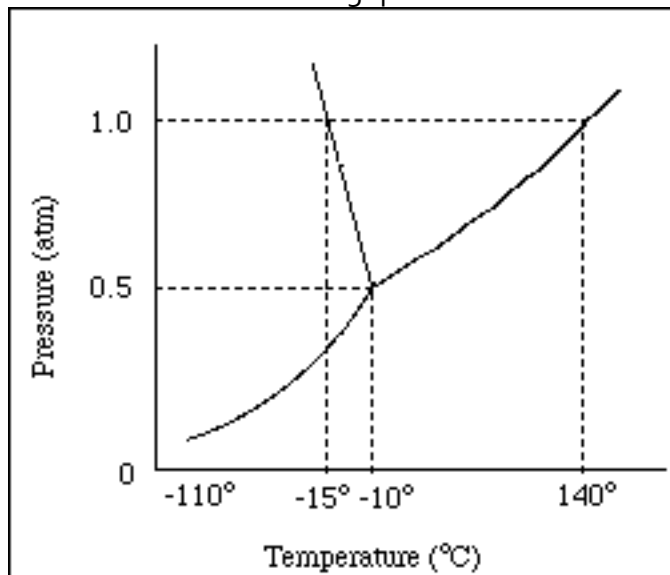
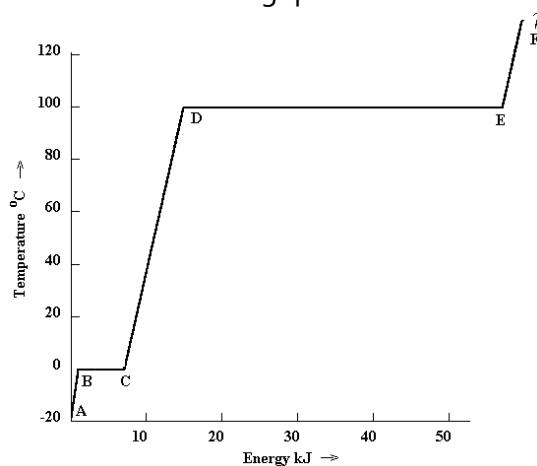


Ch. 9 Cramsheet!

1. Define intermolecular force:
2. What are the three types of intermolecular forces?
 - a. Which one is the strongest?
 - b. Which one is present in all molecules?
 - c. Which is the weakest?
 - d. Which is present in ALL polar molecules?
3. Use the phase diagram below to answer the following questions:



- a. Label the three phases: solid, liquid and gas
 - b. What line represents the melting points?
 - c. What line represents the boiling points?
 - d. Label the triple point and critical point. What is the significance of each?
 - e. What does the term "normal" mean for this phase diagram?
 - f. What is the normal melting and boiling point of this substance?
4. Use the heating curve below to answer the following questions:



- a. Between what points is boiling occurring? Melting?
- b. At what point is the kinetic energy the highest?
- c. The liquid phase is present between what points?
- d. What is the boiling point? What substance is this the heating curve for?