

Name _____ Per. _____ Date _____

Predicting Reactions (Adapted from Paul Groves, SPHS)**SUBSTANCES**

Write the chemical symbol for the following formulas.

If the substance exists mainly as separate ions, write the separate ions.

solid aluminum oxide	
a solution of sodium hydroxide	
solid calcium oxide	
sulfur trioxide gas	
0.1-molar sulfuric acid	
0.1-molar potassium hydroxide	
calcium metal	
nitrogen gas	
solid copper(II) sulfide	
oxygen gas	
concentrated solution of hydrochloric acid	
powdered manganese dioxide	
concentrated solution of ammonia	
a solution of zinc iodide	
a solution of copper(II) sulfate	
a solution of barium hydroxide	
a solution of magnesium nitrate	
solid lithium hydride	
water	
a solution of ammonia	
a solution of hydrofluoric acid	
a piece of aluminum metal	
a solution of silver nitrate	
a solution of potassium iodide	
solid potassium oxide	
an excess of nitric acid solution	
carbon dioxide gas	

a suspension of calcium carbonate	
a strip of copper	
dilute nitric acid	
potassium permanganate solution	
an acidic solution of hydrogen peroxide	
solid manganese (II) sulfide	
chlorine gas	
hot iron filings	
solid magnesium nitride	
sulfur dioxide gas	
a suspension of silver chloride	
a solution of tri-potassium phosphate	
a solution of zinc nitrate	
sodium cyanide solution	
a solution of manganese(II) sulfate	
a solution of ammonium sulfide	
phosphorus(V) oxide powder	
solid ammonium carbonate	
solid potassium permanganate	
a small piece of sodium metal	
a solution of potassium dichromate	
an acidified solution of iron(II) chloride	
solid barium oxide	
a solution of iron(II) nitrate	
solid calcium phosphate	
hydrogen sulfide gas	
a solution of mercury(II) chloride	