d. sulfurous acid

H_2SO_3

98. Write the formula for each of the following.

a. silicon dioxide

SiO_2

b. bromous acid

$HBrO_2$

c. chlorine trifluoride

ClF_3

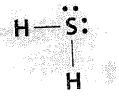
d. hydrobromic acid

HBr

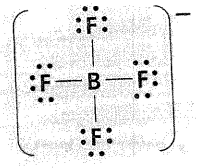
Lewis Structures (9.3)

99. Draw the Lewis structure for each of these molecules or ions.

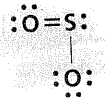
a. H_2S



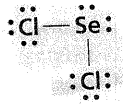
b. BF_4^-



c. SO_2

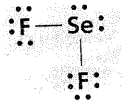


d. $SeCl_2$

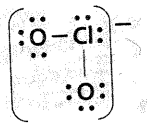


100. Draw the Lewis structure for each of these molecules or ions.

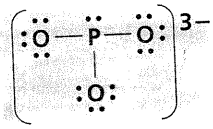
a. SeF_2



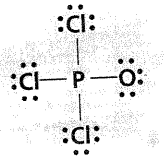
b. ClO_2^-



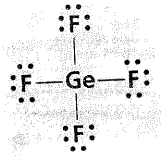
c. PO_3^{3-}



d. $POCl_3$



e. GeF_4



101. Which of the following elements are capable of forming molecules in which an atom has an expanded octet? Explain your answer.

a. B
b. C
c. P
d. O
e. Se

P and Se because they are period 3 and higher and have a d sublevel available