

## Kinetics - Chapter 17

I.

A. Chemical \_\_\_\_\_ :

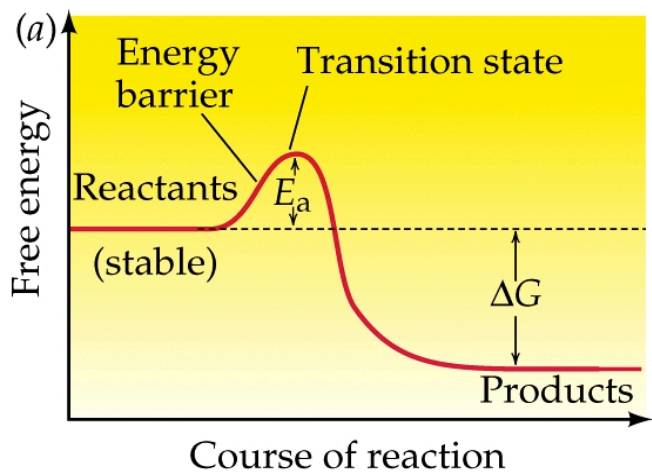
B. \_\_\_\_\_ :

1. Over time, the \_\_\_\_\_ and the \_\_\_\_\_
2. Reaction rate varies with the type of reaction
  - a. combustion = \_\_\_\_\_
  - b. rusting = \_\_\_\_\_

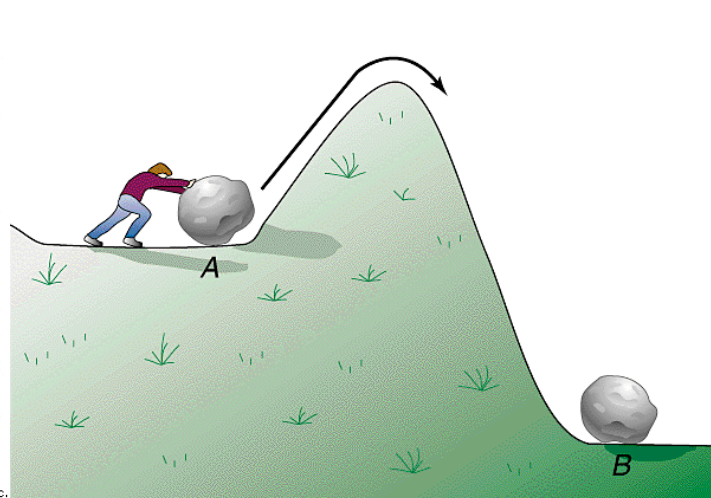
## II. The Reaction Process

A. \_\_\_\_\_ :

1. Theory states that reactants must \_\_\_\_\_ for a reaction to occur.
2. BUT,
3. In order for an \_\_\_\_\_ (creation of new products) to occur molecules \_\_\_\_\_ with:
  - a. \_\_\_\_\_ (known as \_\_\_\_\_).
  - b. the \_\_\_\_\_
4. The \_\_\_\_\_
5. The minimum amount of \_\_\_\_\_ is called \_\_\_\_\_

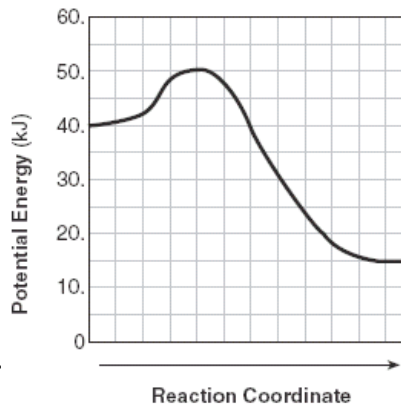
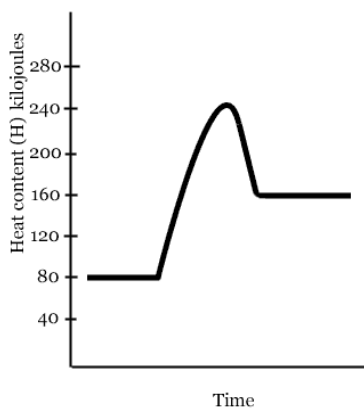


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Is this an endothermic or exothermic reaction?  
: shows energy changes as a reaction proceeds.

6.
  - a. When the \_\_\_\_\_
  - b. When the \_\_\_\_\_
  - c. The activation energy of any reaction is the \_\_\_\_\_



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## III. Factors that Affect Reaction Rates

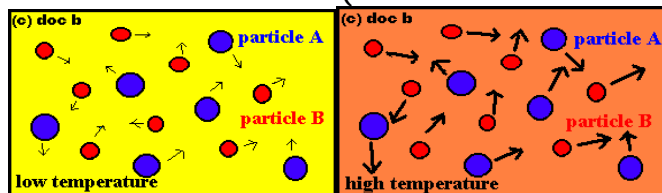
: how complex and reactive the reactants are.

A.

B. :

1. as the
2. This is due to an
3. Moving =

(two birds with one stone!)



C.

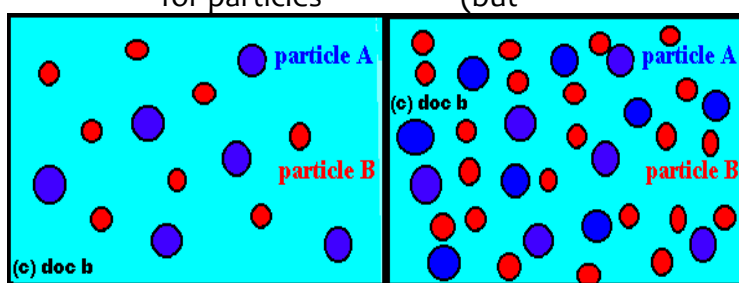
:

1. as the
2. =

for particles

(but

)!



D.

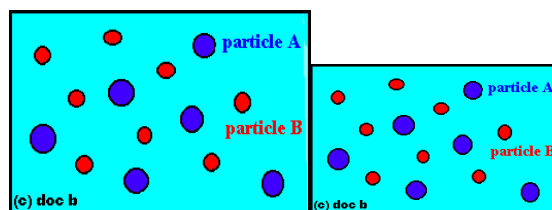
:

1. Rate increases as the pressure of reactants increases.
2. When pressure increases,
3. If you
4. This is only true for

=

(again

)!



E.

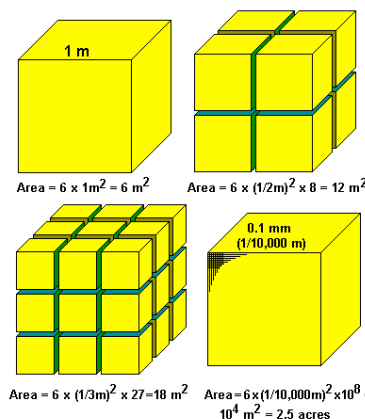
:

1. An
- 2.

will

.

=



F.

:

- 1.
2. in the reaction
3. Catalysts by
4. In order to do this, the catalysts create a different pathway.
- e. are the body's catalysts!

