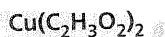
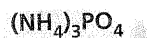


4. Determine the formula for ionic compounds composed of the following ions.

a. copper(II) and acetate



b. ammonium and phosphate



c. calcium and hydroxide



d. gold(III) and cyanide



5. Name the following compounds.

a. $\text{Co}(\text{OH})_2$

cobalt(II) hydroxide

b. $\text{Ca}(\text{ClO}_3)_2$

calcium chlorate

c. Na_3PO_4

sodium phosphate

d. $\text{K}_2\text{Cr}_2\text{O}_7$

potassium dichromate

e. SrI_2

strontium iodide

f. HgF_2

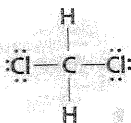
mercury(II) fluoride

Chapter 9

Section 9-1

1. Draw the Lewis structure for the following molecules.

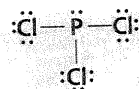
a. CCl_2H_2



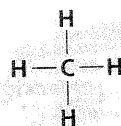
b. HF



c. PCl_3



d. CH_4



Section 9-2

2. Name the following binary compounds.

a. S_4N_2

tetrasulfur dinitride

b. OCl_2

oxygen dichloride

c. SF_6

sulfur hexafluoride

d. NO

nitrogen monoxide

e. SiO_2

silicon dioxide

f. IF_7

iodine heptafluoride

3. Name the following acids: H_3PO_4 , HBr, HNO_3 .

phosphoric acid; hydrobromic acid; nitric acid

Section 9-3

4. Draw the Lewis structure for each of the following.

a. CO



b. CH_2O

