

# Electron Configurations Worksheet

Name \_\_\_\_\_  
Period \_\_\_\_\_  
Date \_\_\_\_\_

Write the complete ground state electron configurations for the following:

1) lithium \_\_\_\_\_

2) oxygen \_\_\_\_\_

3) calcium \_\_\_\_\_

4) titanium \_\_\_\_\_

5) rubidium \_\_\_\_\_

6) lead \_\_\_\_\_

7) erbium \_\_\_\_\_

8) helium \_\_\_\_\_

9) nitrogen \_\_\_\_\_

10) chlorine \_\_\_\_\_

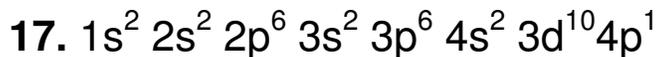
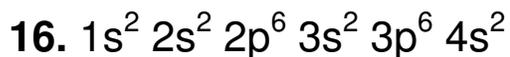
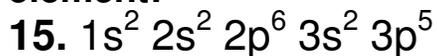
11) iron \_\_\_\_\_

12) zinc \_\_\_\_\_

13) barium \_\_\_\_\_

14) polonium \_\_\_\_\_

**Use the following electron configurations and your periodic table to identify the element:**



**Use the following clues to identify the element. Show any figuring in the space below.**

**18.** This element has a 3p sublevel that contains 3 electrons.

**19.** This element has a 4s sublevel with 2 electrons for its outermost electrons.

**20.** This element has 1 electron in its 3d sublevel.

**21.** This element has 5 electrons in its 5p sublevel

**22.** This element has a completely filled 3p sublevel for its outermost electrons.

**23.** This element has 2 electrons in its 6p sublevel