

Worksheet: Chemical Equations Review

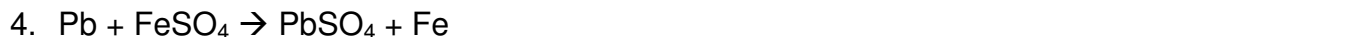
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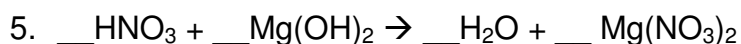
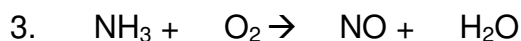
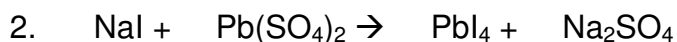
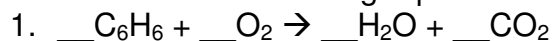
Sections 1: Identify the type of reaction

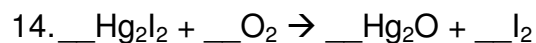
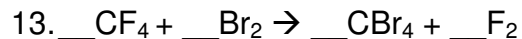
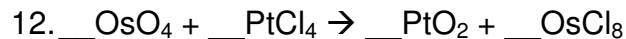
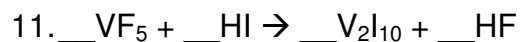
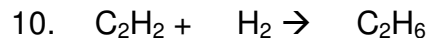
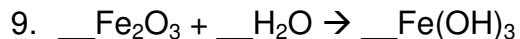
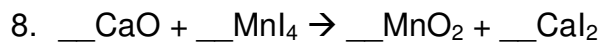
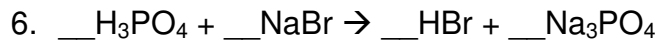
For the following reactions, indicate whether the following are examples of synthesis, decomposition, combustion, single displacement, or double displacement.



Section 2: Practicing Equation Balancing

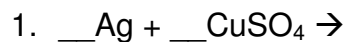
Balance each of the following equations.



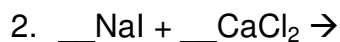


Section 3: Prediction the products of chemical reactions

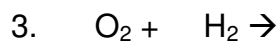
Predict the products of the following reactions, identify the type of reactions, and balance the equation.



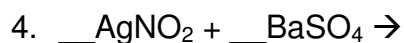
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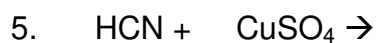
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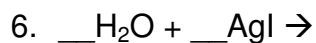
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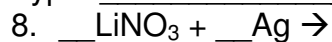
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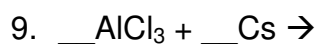
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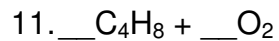
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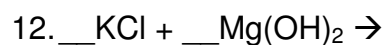
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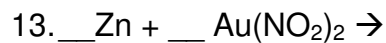
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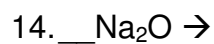
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Section 4 Writing and balancing chemical equations

Write the balanced chemical equation.

1. Solid lithium and liquid water react to form lithium hydroxide in aqueous solution and hydrogen gas.

2. Solid copper (II) sulfide and oxygen gas produce solid copper (II) oxide and sulfur dioxide gas.

3. Sulfur dioxide gas and oxygen gas combine to form sulfur trioxide gas.

4. Water vapor reacts with potassium metal to produce hydrogen and aqueous potassium hydroxide.

5. Solid iron, liquid water, and oxygen gas combine to produce iron (II) hydroxide.

6. Tetraphosphorus decoxide + water \longrightarrow Hydrogen phosphate

7. Sodium hydroxide + Chlorine \longrightarrow Sodium chloride + Sodium hypochlorite + water

8. Potassium chlorate \longrightarrow Potassium perchlorate + Potassium chloride